Gavin R Hedwig

List of Publications by Year
in descending order

29 papers	584 citations	567281 15 h-index	610901 24 g-index
29 all docs	29 docs citations	29 times ranked	189 citing authors

#	Article	IF	Citations
1	Thermodynamic properties of peptide solutions 3. Partial molar volumes and partial molar heat capacities of some tripeptides in aqueous solution. Journal of Solution Chemistry, 1988, 17, 383-397.	1.2	122
2	Thermodynamic properties of peptide solutions: 7. Partial molar isentropic pressure coefficients of some dipeptides in aqueous solution. Journal of Solution Chemistry, 1991, 20, 1113-1127.	1.2	46
3	Group additivity schemes for the calculation of the partial molar heat capacities and volumes of unfolded proteins in aqueous solution. Biophysical Chemistry, 2002, 100, 239-260.	2.8	46
4	Thermodynamic properties of peptide solutions: 14. Partial molar expansibilities and isothermal compressibilities of some glycyl dipeptides in aqueous solution. Journal of Solution Chemistry, 1996, 25, 615-633.	1.2	39
5	The partial molar heat capacities and volumes of some N-acetyl amino acid amides in aqueous solution over the temperature range 288.15 to 328.15 K. Physical Chemistry Chemical Physics, 2000, 2, 1795-1802.	2.8	37
6	Relative partial molar enthalpies and apparent molar volumes of dipeptides in aqueous solution. Journal of Solution Chemistry, 1981, 10, 321-331.	1.2	32
7	Thermodynamic Properties of Peptide Solutions. 16. Partial Molar Heat Capacities and Volumes of Some Tripeptides of Sequence Gly-X-Gly in Aqueous Solution at 25 Â°C. Journal of Chemical & Engineering Data, 1998, 43, 477-481.	1.9	24
8	The partial molar volumes at T=(288.15 to 313.15)K, and the partial molar heat capacities and expansions at T=298.15K of cytidine, uridine, and adenosine in aqueous solution. Journal of Chemical Thermodynamics, 2008, 40, 957-965.	2.0	20
9	Partial molar volumes and heat capacities of single ions in aqueous solution over the temperature range 288.15 to 328.15 KElectronic supplementary information (ESI) available: Densities, apparent molar volumes and apparent molar heat capacities of aqueous solutions of NaBPh4, Ph4PBr, Ph4PCl and NaBr. See http://www.rsc.org/suppdata/cp/b4/b407875f/. Physical Chemistry Chemical Physics, 2004, 6, 4690.	2.8	19
10	Volumetric properties of the glycyl group of proteins in aqueous solution at high pressures. Physical Chemistry Chemical Physics, 2008, 10, 884-897.	2.8	19
11	Partial molar isentropic and isothermal compressibilities of some N-acetyl amino acid amides in aqueous solution at 298.15 K. Physical Chemistry Chemical Physics, 2004, 6, 2440-2445.	2.8	18
12	Isentropic and isothermal compressibilities of the backbone glycyl group of proteins in aqueous solution. Biophysical Chemistry, 2006, 124, 35-42.	2.8	17
13	Volumetric properties of the nucleosides adenosine, cytidine, and uridine in aqueous solution at T=298.15K and p=(10 to 120)MPa. Journal of Chemical Thermodynamics, 2013, 61, 117-125.	2.0	17
14	Title is missing!. Journal of Solution Chemistry, 2001, 30, 861-883.	1.2	16
15	Thermodynamic Properties of Peptide Solutions. Part 17. Partial Molar Volumes and Heat Capacities of the Tripeptides GlyAspGly and GlyGluGly, and Their Salts K[GlyAspGly] and Na[GlyGluGly] in Aqueous Solution at 25 Â°C. Journal of Solution Chemistry, 2005, 34, 801-821.	1.2	16
16	Partial Molar Isentropic and Isothermal Compressions of the Nucleosides Adenosine, Cytidine, and Uridine in Aqueous Solution at 298.15 K. Journal of Chemical & Engineering Data, 2011, 56, 2266-2272.	1.9	16
17	Thermodynamic properties of peptide solutions. Part 15. Partial molar isentropic compressibilities of some glycyl dipeptides in aqueous solution at 15 and 35ï¿½C. Journal of Solution Chemistry, 1996, 25, 1041-1053.	1.2	13
18	Volumetric Properties of Tripeptides with Polar Side-Chains: Partial Molar Volumes at (288.15 to 313.15) K and Partial Molar Expansions at 298.15 K of Some Peptides of Sequence Gly-X-Gly in Aqueous Solution. Journal of Chemical & Engineering Data, 2009, 54, 606-612.	1.9	11

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19	Volumetric interaction coefficients for some nucleosides in aqueous solution at T=298.15K. Journal of Chemical Thermodynamics, 2013, 59, 188-194.	2.0	11
20	Thermodynamic properties of peptide solutions 20. Partial molar volumes and isothermal compressions for some tripeptides of sequence gly-X-gly (X = gly, ala, leu, asn, thr, and tyr) in aqueous solution at T = 298.15 K and p = (10â€"120) MPa. Journal of Chemical Thermodynamics, 2016, 99, 30-39.	2.0	10
21	Thermodynamic Properties of Peptide Solutions. Part 18. Partial Molar Isentropic Compressibilities of Gly-X-Gly Tripeptides (X = Tyr, Pro, Gln, Asp and Glu), and the Peptide Salts K[GlyAspGly], Na[GlyGluGly] and GlyLysGly Acetate in Aqueous Solution at 25 â˜C. Journal of Solution Chemistry, 2005, 34, 1297-1310.	1.2	9
22	The Partial Molar Heat Capacities and Expansions ofÂInosine, 2â€²-Deoxyinosine and 2â€²-Deoxyguanosine inÂAqueous Solution at 298.15ÂK. Journal of Solution Chemistry, 2009, 38, 1315-1331.	1.2	6
23	The Partial Molar Volume and Heat Capacity of the Glycyl Group in Aqueous Solution at 25^C. Journal of Solution Chemistry, 2006, 35, 939-949.	1.2	4
24	Partial Molar Isentropic Compressions of Some Tetra- and Pentapeptides in Aqueous Solution: Implications for Group Additivity Schemes for Unfolded Proteins. Journal of Solution Chemistry, 2012, 41, 690-701.	1.2	3
25	Volumetric Properties at High Pressures of the Nucleosides Inosine, 2â€²-Deoxyinosine, and 2â€²-Deoxyguanosine and the Volumetric Properties of Guanosine Derived Using Group Additivity Methods. Journal of Chemical & Engineering Data, 2014, 59, 3593-3604.	1.9	3
26	Volumetric Properties of the Nucleoside Thymidine in Aqueous Solution at TÂ=Â298.15ÂK and pÂ=Â(10 to 100) MPa. Journal of Solution Chemistry, 2014, 43, 804-820.	1.2	3
27	The Partial Molar Isothermal Compressions of the Nucleosides Adenosine, Cytidine, and Uridine in Aqueous Solution at TÂ=Â(288.15 and 313.15)ÂK. Journal of Solution Chemistry, 2017, 46, 849-861.	1.2	3
28	Volumetric Properties of the Nucleosides Adenosine, Cytidine, and Uridine in Aqueous Solution at Tâ€‰=â€‰(288.15 and 313.15) K and pâ€‰=â€‰(10 to 100) MPa. Journal of Solution Chemistry, 2019, 48, 180-199.	1.2	3
29	The Partial Molar Volumes and Heat Capacities ofÂtheÂArginyl Side-chain of Proteins in Aqueous Solution over the Temperature Range 288.15 toÂ328.15ÂK. Journal of Solution Chemistry, 2010, 39, 1721-1734.	1.2	1