Jaakko I Partanen

List of Publications by Year in descending order

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54	792	471371	642610
papers	citations	17 h-index	23 g-index

54	54	54	263
all docs	docs citations	times ranked	citing authors

#	Article	IF	Citations
1	Re-evaluation of the Activity Coefficients of Aqueous Hydrochloric Acid Solutions up to a Molality of 16.0 molÂ·kgâ¯1 Using the HÃ¼ckel and Pitzer Equations at Temperatures from 0 to 50 Â°C. Journal of Solution Chemistry, 2007, 36, 39-59.	0.6	47
2	Re-Evaluation of the Thermodynamic Activity Quantities in Aqueous Sodium and Potassium Chloride Solutions at 25 Â°C. Journal of Chemical & Engineering Data, 2009, 54, 208-219.	1.0	41
3	Re-Evaluation of the Thermodynamic Activity Quantities in Aqueous Rubidium and Cesium Chloride Solutions at 25 Â°C. Journal of Chemical & Engineering Data, 2010, 55, 249-257.	1.0	31
4	Activity and osmotic coefficients of dilute sodium chloride solutions at 273 K. Journal of Chemical & Engineering Data, 1991, 36, 432-435.	1.0	30
5	Re-Evaluation of the Thermodynamic Activity Quantities in Aqueous Lithium Chloride Solutions at 25 Â°C up to a Molality of 6.0 molÂ·kg‹sup›â¯1‹/sup›. Journal of Chemical & Engineering Data, 2009, 54, 882-889.	1.0	26
6	Re-evaluation of the Thermodynamic Activity Quantities in Aqueous Solutions of Silver Nitrate, Alkali Metal Fluorides and Nitrites, and Dihydrogen Phosphate, Dihydrogen Arsenate, and Thiocyanate Salts with Sodium and Potassium Ions at 25 Â°C. Journal of Chemical & Engineering Data, 2011, 56, 2044-2062.	1.0	26
7	Equations for calculation of the pH of buffer solutions containing sodium or potassium dihydrogen phosphate, sodium hydrogen phosphate, and sodium chloride at 25Â°C. Journal of Solution Chemistry, 1997, 26, 709-727.	0.6	24
8	Determination of Stoichiometric Dissociation Constants of Acetic Acid in Aqueous Solutions Containing Acetic Acid, Sodium Acetate, and Sodium Chloride at (0 to 60) Â°C. Journal of Chemical & Engineering Data, 2003, 48, 797-807.	1.0	24
9	Re-evaluation of the Second Stoichiometric Dissociation Constants of Phosphoric Acid at Temperatures from (0 to 60) Â°C in Aqueous Buffer Solutions with or without NaCl or KCl. 2. Tests and Use of the Resulting HÃ¼ckel Model Equations. Journal of Chemical & Engineering Data, 2005, 50, 2065-2073.	1.0	24
10	Re-evaluation of the Thermodynamic Activity Quantities in Aqueous Alkali Metal Iodide Solutions at 25 Â°C. Journal of Chemical & Engineering Data, 2010, 55, 3708-3719.	1.0	22
11	Re-Evaluation of the Thermodynamic Activity Quantities in Aqueous Alkali Metal Bromide Solutions at 25 Â°C. Journal of Chemical & Engineering Data, 2010, 55, 2202-2213.	1.0	22
12	Re-evaluation of the Thermodynamic Activity Quantities in Aqueous Solutions of Uni-univalent Alkali Metal Salts of Aliphatic Carboxylic Acids and Thallium Acetate at 25 Â°C. Journal of Chemical & Engineering Data, 2011, 56, 4524-4543.	1.0	20
13	Mean Activity Coefficients and Osmotic Coefficients in Aqueous Solutions of Salts of Ammonium Ions with Univalent Anions at 25 Â°C. Journal of Chemical & Engineering Data, 2012, 57, 2654-2666.	1.0	20
14	Comparison of different methods for calculation of the stoichiometric dissociation constant of acetic acid from results of potentiometric titrations at 298.15 K in aqueous sodium or potassium chloride solutions. Fluid Phase Equilibria, 2000, 169, 149-166.	1.4	19
15	Re-evaluation of the Thermodynamic Activity Quantities in Pure Aqueous Solutions of Chlorate, Perchlorate, and Bromate Salts of Lithium, Sodium or Potassium Ions at 298.15 K. Journal of Solution Chemistry, 2012, 41, 271-293.	0.6	19
16	Re-evaluation of the Mean Activity Coefficients of Strontium Chloride in Dilute Aqueous Solutions from (10 to 60) Â°C and at 25 Â°C up to the Saturated Solution Where the Molality Is 3.520 molÂ·kg‹sup›â€“1‹/sup›. Journal of Chemical & Engineering Data, 2013, 58, 2738-2747.	1.0	19
17	Calculation of Stoichiometric Dissociation Constants of Formic, Acetic, Glycolic and Lactic Acids in Dilute Aqueous Potassium, Sodium or Lithium Chloride Solutions at 298.15 K.. Acta Chemica Scandinavica, 1998, 52, 985-994.	0.7	19
18	Re-evaluation of the thermodynamic activity quantities in aqueous alkali metal nitrate solutions at T= 298.15 K. Journal of Chemical Thermodynamics, 2010, 42, 1485-1493.	1.0	18

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19	Equations for the Calculation of the pH of Buffer Solutions Containing Potassium Hydrogen Phthalate, Dipotassium Phthalate, and Potassium Chloride at 298.15 K. Journal of Chemical & Engineering Data, 1997, 42, 805-813.	1.0	17
20	Re-Evaluation of Stoichiometric Dissociation Constants from Electrochemical Cell Data for Propionic andn-Butyric Acids at (0 to 60) Â°C and for Some Other Aliphatic Carboxylic Acids at (18 or) Tj ETQq0 0 0 rgBT /Overlock 10 T 394-406.	1.0	16
21	Traceable Mean Activity Coefficients and Osmotic Coefficients in Aqueous Calcium Chloride Solutions at 25 Â°C up to a Molality of 3.0 molÂ·kg‹sup›â€“1‹/sup›. Journal of Chemical & Engineering Data, 2012, 57, 3247-3257.	1.0	16
22	Re-evaluation of Activity Coefficients in Dilute Aqueous Hydrobromic and Hydriodic Acid Solutions at Temperatures from 0 to 60ÂÂ°C. Journal of Solution Chemistry, 2013, 42, 190-210.	0.6	16
23	Determination of stoichiometric dissociation constants of lactic acid in aqueous salt solutions at 291.15 and at 298.15 K. Fluid Phase Equilibria, 2003, 204, 245-266.	1.4	15
24	Mean Activity Coefficients and Osmotic Coefficients in Dilute Aqueous Sodium or Potassium Chloride Solutions at Temperatures from (0 to 70) Â°C. Journal of Chemical & Engineering Data, 2016, 61, 286-306.	1.0	15
25	Determination of Stoichiometric Dissociation Constants for Propionic Acid in Aqueous Sodium or Potassium Chloride Solutions at 298.15 K.. Acta Chemica Scandinavica, 1999, 53, 547-556.	0.7	15
26	Calculation of the first and second stoichiometric dissociation constants of glycine in aqueous sodium chloride solutions at 298.15 K. Zeitschrift Fur Elektrotechnik Und Elektrochemie, 1998, 102, 855-865.	0.9	14
27	Traceable Thermodynamic Quantities for Dilute Aqueous Sodium Chloride Solutions at Temperatures from (0 to 80) Â°C. Part 1. Activity Coefficient, Osmotic Coefficient, and the Quantities Associated with the Partial Molar Enthalpy. Journal of Chemical & Engineering Data, 2017, 62, 2617-2632.	1.0	14
28	Title is missing!. Journal of Solution Chemistry, 2002, 31, 197-210.	0.6	13
29	Re-evaluation of the First and Second Stoichiometric Dissociation Constants of Phthalic Acid at Temperatures from (0 to 60) Â°C in Aqueous Phthalate Buffer Solutions with or without Potassium Chloride. 1. Estimation of the Parameters for the HÃ¼ckel Model Activity Coefficient Equations for Calculation of the Second Dissociation Constant. Journal of Chemical & Engineering Data, 2006, 51, 777-784.	1.0	13
30	Determination of the Pitzer Interaction Parameters at 273.15 K from the Freezing-Point Data Available for NaCl and KCl Solutions. Industrial & Engineering Chemistry Research, 2014, 53, 5608-5616.	1.8	13
31	Determination of Stoichiometric Dissociation Constants of Formic Acid in Aqueous Sodium or Potassium Chloride Solutions at 298.15 K. Journal of Chemical & Engineering Data, 2000, 45, 110-115.	1.0	12
32	Determination of the Glass Electrode Parameters by Means of Potentiometric Titration of Acetic Acid in Aqueous Sodium or Potassium Chloride Solutions at 25Â°C. Journal of Solution Chemistry, 1999, 28, 413-433.	0.6	11
33	Determination of stoichiometric dissociation constant of n-butyric acid in aqueous sodium or potassium chloride solutions at 298.15 K. Fluid Phase Equilibria, 2000, 173, 135-148.	1.4	11
34	Re-evaluation of the Second Stoichiometric Dissociation Constants of Phosphoric Acid at Temperatures from (0 to 60) Â°C in Aqueous Buffer Solutions with or without NaCl or KCl. 1. Estimation of the Parameters for the HÃ¼ckel Model Activity Coefficient Equations. Journal of Chemical & Engineering Data, 2005, 50, 1502-1509.	1.0	11
35	Activity Coefficients of Potassium Dihydrogen Phosphate in Aqueous Solutions at 25Â°C and in Aqueous Mixtures of Urea and this Electrolyte in the Temperature Range 20â€“35Â°C. Zeitschrift Fur Physikalische Chemie, 2003, 217, 723-738.	1.4	10
36	Re-evaluation of the First and Second Stoichiometric Dissociation Constants of Phthalic Acid at Temperatures from (0 to 60) Â°C in Aqueous Phthalate Buffer Solutions with or without Potassium Chloride. 2. Estimation of Parameters for the Model for the First Dissociation Constant and Tests and Use of the Resulting Activity Coefficient Equations. Journal of Chemical & Engineering Data, 2006, 51, 2065-2073.	1.0	10

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37	Traceable mean activity coefficients and osmotic coefficients in aqueous magnesium chloride solutions at T=298.15K up to a molality of 3.0molÂ·kgâ¯1. Journal of Chemical Thermodynamics, 2013, 66, 65-70.	1.0	10
38	Title is missing!. Journal of Solution Chemistry, 2002, 31, 187-196.	0.6	9
39	Traceable Values for Activity and Osmotic Coefficients in Aqueous Sodium Chloride Solutions at Temperatures from 273.15 to 373.15 K up to the Saturated Solutions. Journal of Chemical & Engineering Data, 2020, 65, 5226-5239.	1.0	9
40	Traceable mean activity coefficients of barium chloride in dilute aqueous solutions from T=(273 to) Tj ETQq0 0 0 rgBT /Overlock 10 Tf 50 Chemical Thermodynamics, 2014, 75, 128-135.	1.0	8
41	Re-evaluation of Stoichiometric Dissociation Constants from Electrochemical Cell Data for Formic Acid at Temperatures from (0 to 60) Â°C and for Some Other Aliphatic Carboxylic Acids at (18 or 25) Â°C in Aqueous Potassium Chloride Solutions. Journal of Chemical & Engineering Data, 2005, 50, 497-507.	1.0	7
42	Traceable Thermodynamic Quantities for Dilute Aqueous Sodium Chloride Solutions at Temperatures from (0 to 80) Â°C. Part 2. The Quantities Associated with the Partial Molar Heat Capacity. Journal of Chemical & Engineering Data, 2017, 62, 4215-4227.	1.0	7
43	Title is missing!. Journal of Solution Chemistry, 2001, 30, 443-462.	0.6	6
44	Re-evaluation of the second thermodynamic dissociation constants of Î±-alanine, valine, and leucine using potentiometric data measured for aqueous potassium chloride solutions at 298.15 K. Canadian Journal of Chemistry, 2005, 83, 46-56.	0.6	6
45	Re-evaluation of the First and Second Stoichiometric Dissociation Constants of Oxalic Acid at Temperatures from 0 to 60â€‰Â°C in Aqueous Oxalate Buffer Solutions withÂorÂwithout Sodium or Potassium Chloride. Journal of Solution Chemistry, 2009, 38, 1385-1416.	0.6	6
46	Determination of the second stoichiometric dissociation constants of glycine in aqueous sodium or potassium chloride solutions at 298.15 K. Canadian Journal of Chemistry, 2003, 81, 1462-1470.	0.6	5
47	Traceable Activity and Osmotic Coefficients in Pure Aqueous Solutions of Alkaline Earth Metal Bromides and Iodides at 25 Â°C. Journal of Chemical & Engineering Data, 2014, 59, 2530-2540.	1.0	5
48	Traceable Thermodynamic Quantities for Dilute Aqueous NaCl Solutions at Temperatures from (353.15) Tj ETQq0 0 0 rgBT /Overlock 10 Engineering Data, 2019, 64, 16-33.	1.0	5
49	Thermodynamically Traceable Calorimetric Results for Dilute Aqueous Potassium Chloride Solutions at Temperatures from 273.15 to 373.15 K. Part 2. The Quantities Associated with the Partial Molar Heat Capacity. Journal of Chemical & Engineering Data, 2019, 64, 3971-3982.	1.0	4
50	Thermodynamically Traceable Calorimetric Results for Dilute Aqueous Potassium Chloride Solutions at Temperatures from (273.15 to 373.15) K. Part 1. The Quantities Associated with the Partial Molar Enthalpy. Journal of Chemical & Engineering Data, 2019, 64, 2519-2535.	1.0	4
51	Determination of the Pitzer Interaction Parameters at 273.15 K from the Freezing-Point Data Available for Solutions of Uni-Univalent Electrolytes. Industrial & Engineering Chemistry Research, 2014, 53, 19351-19358.	1.8	3
52	Determination of Stoichiometric Dissociation Constants of Glycolic Acid in Dilute Aqueous Sodium or Potassium Chloride Solutions at 298.15 K. Zeitschrift Fur Physikalische Chemie, 2001, 215, .	1.4	2
53	pH Standardization of 0.05 molÂ·kg-1Tetraoxalate Buffer at Temperatures from (5 to 45) Â°C with Added KCl Molality up to 1.0 molÂ·kg-1. Journal of Chemical & Engineering Data, 2007, 52, 973-976.	1.0	2
54	Determination of Stoichiometric Dissociation Constant of Ammonium Ion in Aqueous Potassium Chloride Solutions at 298.15Â K. Zeitschrift Fur Physikalische Chemie, 2005, 219, 1609-1623.	1.4	1