

Dan Fraenkel

List of Publications by Year in descending order

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23
papers

406
citations

758635

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h-index

752256

20
g-index

25
all docs

25
docs citations

25
times ranked

202
citing authors

#	ARTICLE	IF	CITATIONS
1	The electric conductance of dilute sulfuric acid in water: a new theoretical interpretation. <i>Molecular Physics</i> , 2021, 119, e1905898.	0.8	3
2	Recommended Activity Coefficients and Equivalent Conductivities of Aqueous Hydrochloric Acid in the Concentration Range 5×10^{-6} – 10^{-1} M and Temperature Range 5 – 55 °C Based on the DH α S and DHO α S Theoretical Treatments. <i>Journal of Chemical & Engineering Data</i> , 2021, 66, 2470-2479.	1.0	3
3	Correlation between the specific conductivity and the equivalent conductivity of an individual ion in electrolyte solution. <i>Chemical Physics Letters</i> , 2021, 781, 138957.	1.2	6
4	Interconversion of Specific and Equivalent Conductivity of Ions in Electrolyte Solution: Effects of High Ionic Valence and Temperature. <i>Journal of Chemical Theory and Computation</i> , 2021, , .	2.3	0
5	Negative Deviations from the Debye-Hückel Limiting Law for High-Charge Polyvalent Electrolytes: Are They Real?. <i>Journal of Chemical Theory and Computation</i> , 2018, 14, 2609-2620.	2.3	9
6	An improved theory of the electric conductance of ionic solutions based on the concept of the ion-atmosphere's smaller-ion shell. <i>Physical Chemistry Chemical Physics</i> , 2018, 20, 29896-29909.	1.3	15
7	A new theoretical development of the limiting electric conductivity of ions in solution. <i>Molecular Physics</i> , 2018, 116, 2271-2293.	0.8	10
8	Theoretical interpretation of the limiting electric conductivity in ionic solution. <i>Molecular Physics</i> , 2017, 115, 2944-2950.	0.8	2
9	Ion strength limit of computed excess functions based on the linearized Poirsson-Boltzmann equation. <i>Journal of Computational Chemistry</i> , 2015, 36, 2302-2316.	1.5	4
10	Structure and ionization of sulfuric acid in water. <i>New Journal of Chemistry</i> , 2015, 39, 5124-5136.	1.4	22
11	Computing Excess Functions of Ionic Solutions: The Smaller-Ion Shell Model <i>versus</i> the Primitive Model. 1. Activity Coefficients. <i>Journal of Chemical Theory and Computation</i> , 2015, 11, 178-192.	2.3	16
12	Computing Excess Functions of Ionic Solutions: The Smaller-Ion Shell Model <i>versus</i> the Primitive Model. 2. Ion-Size Parameters. <i>Journal of Chemical Theory and Computation</i> , 2015, 11, 193-204.	2.3	13
13	Agreement of electrolyte models with activity coefficient data of sulfuric acid in water. <i>Journal of Chemical Thermodynamics</i> , 2014, 78, 215-224.	1.0	12
14	Theoretical analysis of aqueous solutions of mixed strong electrolytes by a smaller-ion shell electrostatic model. <i>Journal of Chemical Physics</i> , 2014, 140, 054513.	1.2	11
15	Electrolytic Nature of Aqueous Sulfuric Acid. 1. Activity. <i>Journal of Physical Chemistry B</i> , 2012, 116, 11662-11677.	1.2	21
16	Single-Ion Activity: Experiment versus Theory. <i>Journal of Physical Chemistry B</i> , 2012, 116, 3603-3612.	1.2	46
17	Electrolytic Nature of Aqueous Sulfuric Acid. 2. Acidity. <i>Journal of Physical Chemistry B</i> , 2012, 116, 11678-11686.	1.2	16
18	Reply to α Comment on 'Single-Ion Activity: Experiment versus Theory'. <i>Journal of Physical Chemistry B</i> , 2012, 116, 13292-13293.	1.2	7

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19	Monoprotic Mineral Acids Analyzed by the Smaller-Ion Shell Model of Strong Electrolyte Solutions. Journal of Physical Chemistry B, 2011, 115, 557-568.	1.2	46
20	Effect of Solvent Permittivity on the Thermodynamic Behavior of HCl Solutions: Analysis Using the Smaller-Ion Shell Model of Strong Electrolytes. Journal of Physical Chemistry B, 2011, 115, 14634-14647.	1.2	23
21	Comment on "The nonmonotonic concentration dependence of the mean activity coefficient of electrolytes is a result of a balance between solvation and ion-ion correlations" [J. Chem. Phys. 133, 154507 (2010)]. Journal of Chemical Physics, 2011, 134, 157101.	1.2	7
22	Acid strength of solids probed by catalytic isobutane conversion. Journal of Catalysis, 2010, 274, 29-51.	3.1	21
23	Simplified electrostatic model for the thermodynamic excess potentials of binary strong electrolyte solutions with size-dissimilar ions. Molecular Physics, 2010, 108, 1435-1466.	0.8	89