

# Tae Hoon Choi

## List of Publications by Year in descending order

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Version: 2024-02-01

9  
papers

104  
citations

1684188  
5  
h-index

1474206  
9  
g-index

9  
all docs

9  
docs citations

9  
times ranked

180  
citing authors

#	ARTICLE	IF	CITATIONS
1	Vibrational Signatures of $\text{HNO}_3$ Acidity When Complexed with Microhydrated Alkali Metal Ions, $M^+(\text{HNO}_3)(\text{H}_2\text{O})_n$ ( $M = \text{Li, K, Na, Rb, Cs}$ ), at 20 K. <i>Journal of Physical Chemistry A</i> , 2022, 126, 1640-1647.	2.5	4
2	Water Network Shape-Dependence of Local Interactions with the Microhydrated $\text{NO}_2^-$ and $\text{CO}_2^-$ Anionic Head Groups by Cold Ion Vibrational Spectroscopy. <i>Journal of Physical Chemistry A</i> , 2022, 126, 2471-2479.	2.5	2
3	Mapping the temperature-dependent and network site-specific onset of spectral diffusion at the surface of a water cluster cage. <i>Proceedings of the National Academy of Sciences of the United States of America</i> , 2020, 117, 26047-26052.	7.1	15
4	Molecular-level origin of the carboxylate head group response to divalent metal ion complexation at the air-water interface. <i>Proceedings of the National Academy of Sciences of the United States of America</i> , 2019, 116, 14874-14880.	7.1	37
5	Model potential study of non-valence correlation-bound anions of $(\text{C}_{60})_n^-$ clusters: the role of electric field-induced charge transfer. <i>Faraday Discussions</i> , 2019, 217, 547-560.	3.2	8
6	Conformers of Zwitterionic Glycine in Aqueous Phase. <i>Bulletin of the Korean Chemical Society</i> , 2018, 39, 227-230.	1.9	1
7	Implementation of analytical gradients and of a mixed real and momentum space DVR method for excess electron systems described by a self-consistent polarization model. <i>Journal of Chemical Physics</i> , 2017, 147, 161717.	3.0	2
8	Discrete Variable Representation Implementation of the One-Electron Polarization Model. <i>Journal of Chemical Theory and Computation</i> , 2010, 6, 2388-2394.	5.3	5
9	Application of the SCC-DFTB Method to $\text{H}^+(\text{H}_2\text{O})_6$ , $\text{H}^+(\text{H}_2\text{O})_{21}$ , and $\text{H}^+(\text{H}_2\text{O})_{22}$ . <i>Journal of Physical Chemistry B</i> , 2010, 114, 6932-6936.	2.6	30