

# Peter Stange

## List of Publications by Year in descending order

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28  
papers

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citations

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#	ARTICLE	IF	CITATIONS
1	Spectroscopic Evidence for Clusters of Like-Charged Ions in Ionic Liquids Stabilized by Cooperative Hydrogen Bonding. <i>ChemPhysChem</i> , 2016, 17, 458-462.	2.1	115
2	Ion Speciation of Protic Ionic Liquids in Water: Transition from Contact to Solvent-Separated Ion Pairs. <i>Angewandte Chemie - International Edition</i> , 2013, 52, 2990-2994.	13.8	89
3	Controlling the Subtle Energy Balance in Protic Ionic Liquids: Dispersion Forces Compete with Hydrogen Bonds. <i>Angewandte Chemie - International Edition</i> , 2015, 54, 2792-2795.	13.8	78
4	Equilibrium of Contact and Solvent-Separated Ion Pairs in Mixtures of Protic Ionic Liquids and Molecular Solvents Controlled by Polarity. <i>Angewandte Chemie - International Edition</i> , 2013, 52, 12439-12442.	13.8	59
5	Dispersion and Hydrogen Bonding Rule: Why the Vaporization Enthalpies of Aprotic Ionic Liquids Are Significantly Larger than those of Protic Ionic liquids. <i>Angewandte Chemie - International Edition</i> , 2016, 55, 11682-11686.	13.8	50
6	Hydrogen Bonding Between Ions of Like Charge in Ionic Liquids Characterized by NMR Deuteron Quadrupole Coupling Constants—Comparison with Salt Bridges and Molecular Systems. <i>Angewandte Chemie - International Edition</i> , 2019, 58, 17863-17871.	13.8	41
7	Ion Pairing in Protic Ionic Liquids Probed by Far-Infrared Spectroscopy: Effects of Solvent Polarity and Temperature. <i>ChemPhysChem</i> , 2014, 15, 2604-2609.	2.1	40
8	The Double-Faced Nature of Hydrogen Bonding in Hydroxy-Functionalized Ionic Liquids Shown by Neutron Diffraction and Molecular Dynamics Simulations. <i>Angewandte Chemie - International Edition</i> , 2019, 58, 12887-12892.	13.8	40
9	Characterization of Doubly Ionic Hydrogen Bonds in Protic Ionic Liquids by NMR Deuteron Quadrupole Coupling Constants: Differences to H-bonds in Amides, Peptides, and Proteins. <i>Angewandte Chemie - International Edition</i> , 2017, 56, 14310-14314.	13.8	35
10	Like-Likes-Like: Cooperative Hydrogen Bonding Overcomes Coulomb Repulsion in Cationic Clusters with Net Charges up to $Q = +6$ e. <i>ChemPhysChem</i> , 2018, 19, 1691-1695.	2.1	29
11	When hydrogen bonding overcomes Coulomb repulsion: from kinetic to thermodynamic stability of cationic dimers. <i>Physical Chemistry Chemical Physics</i> , 2019, 21, 8215-8220.	2.8	29
12	Controlling “like-likes”-like-charge attraction in hydroxy-functionalized ionic liquids by polarizability of the cations, interaction strength of the anions and varying alkyl chain length. <i>Physical Chemistry Chemical Physics</i> , 2020, 22, 2763-2774.	2.8	29
13	Dynamical heterogeneities in ionic liquids as revealed from deuteron NMR. <i>Chemical Communications</i> , 2018, 54, 3098-3101.	4.1	21
14	Dissecting the Vaporization Enthalpies of Ionic Liquids by Exclusively Experimental Methods: Coulomb Interaction, Hydrogen Bonding, and Dispersion Forces. <i>Angewandte Chemie - International Edition</i> , 2019, 58, 8589-8592.	13.8	19
15	Counting cations involved in cationic clusters of hydroxy-functionalized ionic liquids by means of infrared and solid-state NMR spectroscopy. <i>Physical Chemistry Chemical Physics</i> , 2020, 22, 6861-6867.	2.8	17
16	Steuerung der subtilen Energiebalance in protischen ionischen Flüssigkeiten: Dispersionskräfte im Wettstreit mit Wasserstoffbrücken. <i>Angewandte Chemie</i> , 2015, 127, 2834-2837.	2.0	13
17	Dissecting intermolecular interactions in the condensed phase of ibuprofen and related compounds: the specific role and quantification of hydrogen bonding and dispersion forces. <i>Physical Chemistry Chemical Physics</i> , 2020, 22, 4896-4904.	2.8	13
18	Dispersion und Wasserstoffbrücken bestimmend – Warum die Verdampfungsenthalpien von aprotischen größer als die von protischen ionischen Flüssigkeiten sind. <i>Angewandte Chemie</i> , 2016, 128, 11856-11860.	2.0	10

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19	Freezing the Motion in Hydroxy-Functionalized Ionic Liquidsâ€“Temperature Dependent NMR Deuteron Quadrupole Coupling Constants for Two Types of Hydrogen Bonds Far below the Glass Transition. <i>Journal of Physical Chemistry Letters</i> , 2020, 11, 6000-6006.	4.6	10
20	Balance Between Contact and Solvent-Separated Ion Pairs in Mixtures of the Protic Ionic Liquid [Et <sub>3</sub> NH][MeSO <sub>3</sub> ] with Water Controlled by Water Content and Temperature. <i>Journal of Physical Chemistry B</i> , 2021, 125, 4476-4488.	2.6	9
21	Simultaneous determination of deuteron quadrupole coupling constants and rotational correlation times: the model case of hydrogen bonded ionic liquids. <i>Physical Chemistry Chemical Physics</i> , 2019, 21, 25597-25605.	2.8	8
22	Structure, hydrogen bond dynamics and phase transition in a model ionic liquid electrolyte. <i>Physical Chemistry Chemical Physics</i> , 2022, 24, 6064-6071.	2.8	8
23	Hydrogen Bonding Between Ions of Like Charge in Ionic Liquids Characterized by NMR Deuteron Quadrupole Coupling Constantsâ€“Comparison with Salt Bridges and Molecular Systems. <i>Angewandte Chemie</i> , 2019, 131, 18027-18035.	2.0	7
24	Charakterisierung von Wasserstoffbrücken zwischen Ionen in protischen ionischen Flüssigkeiten mittels NMRâ€“Deuteronâ€“Quadrupolâ€“Kopplungskonstanten â€“ Unterschiede zu Hâ€“Brücken in Amiden, Peptiden und Proteinen. <i>Angewandte Chemie</i> , 2017, 129, 14500-14505.	2.0	5
25	Die zweigesichtige Natur der Wasserstoffbrückenbindung in hydroxyfunktionalisierten ionischen Flüssigkeiten, offenbart durch Neutronendiffraktometrie und Molekulardynamikâ€“Simulation. <i>Angewandte Chemie</i> , 2019, 131, 13019-13024.	2.0	5
26	Three in One: The Versatility of Hydrogen Bonding Interaction in Halide Salts with Hydroxyâ€“Functionalized Pyridinium Cations. <i>ChemPhysChem</i> , 2021, 22, 1850-1856.	2.1	5
27	Zerlegung der Verdampfungsenthalpien ionischer Flüssigkeiten durch rein experimentelle Methoden: Coulombâ€“Wechselwirkung, Wasserstoffbrücken und Dispersionskräfte. <i>Angewandte Chemie</i> , 2019, 131, 8679-8683.	2.0	1
28	Zerlegung der Verdampfungsenthalpien ionischer Flüssigkeiten durch rein experimentelle Methoden: Coulombâ€“Wechselwirkung, Wasserstoffbrücken und Dispersionskräfte. <i>Angewandte Chemie</i> , 2019, 131, 8685.	2.0	0