Christomir Christov

List of Publications by Year
in descending order

55 papers	1,317 citations	304743 22 h-index	377865 34 g-index
55 all docs	55 docs citations	55 times ranked	349 citing authors

#	Article	IF	Citations
1	Chemical and geochemical modeling. Thermodynamic models for binary fluoride systems from low to very high concentration (> 35 m) at 298.15 K. Acta Scientifica Naturalis, 2021, 8, 1-15.	0.1	3
2	Modeling the Solution Properties and Mineralâ€“Solution Equilibria in Radionuclide-Bearing Aqueous Nitrate Systems: Application to Binary and Ternary Systems Containing U, Th, or Lanthanides at 25 Â°C. Journal of Chemical & Engineering Data, 2020, 65, 3613-3626.	1.9	7
3	Modeling the Osmotic and Activity Coefficients of Lanthanide Nitrate Aqueous Solutions at 298.15 K from Low Molalities to Supersaturation. Journal of Chemical & Engineering Data, 2019, 64, 345-359.	1.9	12
4	A thermodynamic model for solution behavior and solid-liquid equilibrium in Na-K-Mg-Ca-Al(III)-Fe(III)-Cr(III)-Cl-H₂O system from low to very high concentration at 25Â°C. Acta Scientifica Naturalis, 2019, 6, 26-36.	0.1	2
5	A Pitzer Parametrization To Predict Solution Properties and Salt Solubility in the Hâ€“Naâ€“Kâ€“Caâ€“Mgâ€“NO₃â€“H₂O System at 298.15 K. Journal of Chemical & Engineering Data, 2018, 63, 787-800.	1.9	26
6	Thermodynamic model for solution behavior and solid-liquid equilibrium in Na-Al(III)-Fe(III)-Cr(III)-Cl-H₂O system at 25Â°C. Acta Scientifica Naturalis, 2018, 5, 6-16.	0.1	1
7	Study of bromide salts solubility in the (m1KBr+m2CaBr2)(aq) system at T=323.15K. Thermodynamic model of solution behaviour and (solid+liquid) equilibria in the ternaries (m1KBr+m2CaBr2)(aq), and (m1MgBr2+m2CaBr2)(aq), and in the quinary (Na+K+Mg+Ca+Br+H2O) systems to high concentration and temperature. Journal of Chemical Thermodynamics, 2012, 55, 7-22.	2.0	56
8	Temperature variable chemical model of bromideâ€“sulfate solution interaction parameters and solidâ€“liquid equilibria in the Naâ€“Kâ€“Caâ€“Brâ€“SO4â€“H2O system. Calphad: Computer Coupling of Phase Diagrams and Thermochemistry, 2012, 36, 71-81.	1.6	48
9	Study of bromide salts solubility in the (m1NaBr +m2MgBr2)(aq) system at T= 323.15 K. Thermodynamic model of solution behavior and (solid + liquid) equilibria in the (Na + K + Mg + Br + H2O) system to high concentration and temperature. Journal of Chemical Thermodynamics, 2012, 47, 335-340.	2.0	54
10	Isopiestic investigation of the osmotic coefficients of aqueous CaBr2 and study of bromide salt solubility in the NaBrâ€“CaBr2â€“H2O system at 50ÂÂ° C : Thermodynamic model of solution behavior and solidâ€“liquid equilibria in the CaBr2â€–H2O, and NaBrâ€–CaBr2â€–H2O systems to high concentration and temperature. Calphad: Computer Coupling of Phase Diagrams and Thermochemistry, 2011, 35, 42-53.	1.6	46
11	Isopiestic investigation of the osmotic coefficients of MgBr2(aq) and study of bromide salts solubility in the (m1KBr+m2MgBr2)(aq) system at T=323.15K. Thermodynamic model of solution behaviour and (solid+liquid) equilibria in the MgBr2(aq), and (m1KBr+m2MgBr2)(aq) systems to high concentration and temperature. Journal of Chemical Thermodynamics, 2011, 43, 344-353.	2.0	52
12	Chemical Equilibrium Model of Solution Behavior and Bishofite (MgCl₂Â·6H₂O(cr)) and Hydrogenâ^'Carnallite (HClÂ·MgCl₂Â·7H₂O(cr)) Solubility in the MgCl₂ + H₂O and HClâ^'MgCl₂ + H₂O Systems to High Acid Concentration at (0 to 100) Â°C. Journal of Chemical & Engineering Data, 2009, 54, 2593-2600.	1.9	32
13	Isopiestic Determination of the Osmotic Coefficients of Aqueous MgCl₂ + CaCl₂ Mixed Solution at (25 and 50) Â°C. Chemical Equilibrium Model of Solution Behavior and Solubility in the MgCl₂ + H₂O and MgCl₂ + CaCl₂ + H₂O Systems to High Concentration at (25 and 50) Â°C. Journal of Chemical & Engineering Data, 2009, 54, 627-635.	1.9	49
14	An isopiestic study of aqueous NaBr and KBr at 50Â°C: Chemical equilibrium model of solution behavior and solubility in the NaBrâ€“H2O, KBrâ€“H2O and Naâ€“Kâ€–Brâ€–H2O systems to high concentration and temperature. Geochimica Et Cosmochimica Acta, 2007, 71, 3557-3569.	3.9	79
15	Thermodynamic Modeling of Aqueous Aluminum Chemistry and Solid-Liquid Equilibria to High Solution Concentration and Temperature. I. The Acidic H-Al-Na-K-Cl-H2O System from 0 to 100â€‰Â°C. Journal of Solution Chemistry, 2007, 36, 1495-1523.	1.2	39
16	Thermodynamics of formation of double salts and mixed crystals from aqueous solutions. Journal of Chemical Thermodynamics, 2005, 37, 1036-1060.	2.0	76
17	Thermodynamics of Formation of Double Salts and Mixed Crystals from Aqueous Solutions. ChemInform, 2005, 36, no.	0.0	0
18	Pitzer ion-interaction parameters for Fe(II) and Fe(III) in the quinary {Na+K+Mg+Cl+SO4+H2O} system at T=298.15 K. Journal of Chemical Thermodynamics, 2004, 36, 223-235.	2.0	40

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19	Chemical equilibrium model of solution behavior and solubility in the H-Na-K-OH-Cl-HSO 4 -SO 4 -H 2 O system to high concentration and temperature 1 1Associate editor: D. J. Wesolowski. Geochimica Et Cosmochimica Acta, 2004, 68, 1309-1331.	3.9	107
20	A chemical equilibrium model of solution behavior and solubility in the H-Na-K-Ca-OH-Cl-HSO4-SO4-H2O system to high concentration and temperature. Geochimica Et Cosmochimica Acta, 2004, 68, 3717-3739.	3.9	110
21	Thermodynamics of formation of double salts M2SO4Â·Mâ€²SO4Â·6H2O and M2SeO4Â·Mâ€²SeO4Â·6H2O where M denotes Rb, or Cs and Mâ€² denote Co, Ni, or Zn. Journal of Chemical Thermodynamics, 2003, 35, 1775-1792.	2.0	10
22	Thermodynamic study of aqueous rubidium and cobalt selenate system at the temperature 298.15K. Journal of Chemical Thermodynamics, 2003, 35, 689-697.	2.0	5
23	Thermodynamic study of the aqueous sodium, potassium, and chromium chloride systems at the temperature 298.15K. Journal of Chemical Thermodynamics, 2003, 35, 909-917.	2.0	9
24	Thermodynamic study of the co-crystallization of ammonium, sodium and potassium alums and chromium alums. Calphad: Computer Coupling of Phase Diagrams and Thermochemistry, 2003, 27, 153-160.	1.6	12
25	Thermodynamics of formation of ammonium, sodium and potassium alums and chromium alums. Calphad: Computer Coupling of Phase Diagrams and Thermochemistry, 2002, 26, 85-94.	1.6	28
26	Thermodynamic study of quaternary systems with participation of ammonium and sodium alums and chromium alums. Calphad: Computer Coupling of Phase Diagrams and Thermochemistry, 2002, 26, 341-352.	1.6	28
27	Thermodynamic study of aqueous sodium and potassium chloride and chromate systems at the temperature 298.15 K. Journal of Chemical Thermodynamics, 2002, 34, 987-994.	2.0	13
28	Thermodynamic study of the NaCl î—¸Na2SO4 î—¸Na2Cr2O7 î—¸H2O system at the temperature 298.15 K. Calphad: Computer Coupling of Phase Diagrams and Thermochemistry, 2001, 25, 11-17.	1.6	19
29	Thermodynamic study of the Kî—¸Mgî—¸Alî—¸Clî—¸SO4î—¸H2O system at the temperature 298.15 K. Calphad: Computer Coupling of Phase Diagrams and Thermochemistry, 2001, 25, 445-454.	1.6	33
30	Thermodynamic study of (m1Cs2SeO4+m2NiSeO4)(aq)wheremdenotes molality atT= 298.15 K. Journal of Chemical Thermodynamics, 2001, 33, 1073-1080.	2.0	12
31	Thermodynamic study of the Na-Cu-Cl-SO4- H2O system at the temperature 298.15 K. Journal of Chemical Thermodynamics, 2000, 32, 285-295.	2.0	28
32	Study of (m1LiX+m2CaX2)(aq) wheremidenotes molality andXdenotes Cl, or Br at the temperature 298.15 K. Journal of Chemical Thermodynamics, 2000, 32, 1505-1512.	2.0	34
33	Thermodynamic Study of Na2Cr2O7(aq) at 25 Â°C. Collection of Czechoslovak Chemical Communications, 1999, 64, 595-599.	1.0	10
34	Thermodynamic Study on Aqueous Solutionsof (NH4)2SeO4 and K2SeO4 at 298.15â€‰K. Monatshefte FÃ¼r Chemie, 1999, 130, 1061-1065.	1.8	6
35	Study of (m1KCl+m2MeCl2)(aq) and (m1K2SO4+m2MeSO4)(aq) wheremdenotes molality and Me denotes Cu, or Ni at the temperature 298.15 K. Journal of Chemical Thermodynamics, 1999, 31, 71-83.	2.0	20
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37	Thermodynamic study of (m1Na2SeO4+m2NiSeO4)(aq) wheremdenotes molality at the temperature 298.15 K. Journal of Chemical Thermodynamics, 1998, 30, 73-79.	2.0	6
38	Thermodynamic study of the KCl-K2SO4-K2Cr2O7-H2O system at the temperature 298.15 K. Calphad: Computer Coupling of Phase Diagrams and Thermochemistry, 1998, 22, 449-457.	1.6	24
39	Investigation of the Aqueous Lithium and Nickel Selenate System. Zeitschrift Fur Physikalische Chemie, 1998, 203, 87-93.	2.8	6
40	Thermodynamics of the formation of solid solutions of the type (Me,Meâ€²)SeO4Â·6H2O where Me and Meâ€² denote Mg, Co, Ni, or Zn from aqueous solutions. Journal of Chemical Thermodynamics, 1997, 29, 481-489.	2.0	9
41	Thermodynamic Study of the CuCl2-MCl2-H2O Systems (M = Mg, Co) at 298.15 K. Collection of Czechoslovak Chemical Communications, 1996, 61, 507-511.	1.0	7
42	Thermodynamics of the aqueous sodium and magnesium bromide system at the temperatures 273.15 K and 298.15 K. Calphad: Computer Coupling of Phase Diagrams and Thermochemistry, 1996, 20, 501-509.	1.6	34
43	A Simplified Model for Calculation of the Gibbs Energy of Mixing in Crystals: Thermodynamic Theory, Restrictions and Applicability. Collection of Czechoslovak Chemical Communications, 1996, 61, 1585-1599.	1.0	17
44	Thermodynamics of Formation of Double Salts in the Systems CsClÈ”MCl2â€”H2O where M Denotes Mn, Co, or Cu. Zeitschrift Fur Physikalische Chemie, 1996, 194, 43-50.	2.8	10
45	Thermodynamic study of the co-crystallization of 2RbClÂ·NiCl2Â·2H2O and 2RbClÂ·MnCl2Â·2H2O, at the temperature 298.15 K. Journal of Chemical Thermodynamics, 1996, 28, 743-752.	2.0	9
46	Pitzer Model Based Study of CsX-NiX2-H2O (X = Cl, Br) Systems at 298.15 K. Collection of Czechoslovak Chemical Communications, 1996, 61, 501-506.	1.0	7
47	Isodimorphic co-crystallization of isostructural ammonium chloro- and bromo-carnallites. Journal of Chemical Thermodynamics, 1995, 27, 435-441.	2.0	4
48	Discontinuities in the mixed-crystal series of isostructural carnallite-type double salts. Journal of Chemical Thermodynamics, 1995, 27, 821-828.	2.0	8
49	Thermodynamic study of (b1LiBr +b2MgBr2)(aq), wherebdenotes molality, at the temperature 348.15K. Journal of Chemical Thermodynamics, 1995, 27, 1267-1273.	2.0	26
50	Study of the Conversion of CaSO4 to CaCO3 within the CaSO4 + Na2CO3 = CaCO3 + Na2SO4 Four-Component Water-Salt System. Collection of Czechoslovak Chemical Communications, 1995, 60, 2107-2111.	1.0	9
51	Thermodynamic study of (b1RbCl + b2MeCl2)(aq), where b denotes molality and Me denotes Mn, Co, Ni, or Cu, at the temperature 298.15 K, on the basis of Pitzer's model. Journal of Chemical Thermodynamics, 1994, 26, 1071-1080.	2.0	22
52	Effect of Temperature on the Solubility Diagrams of Ammonium Bromcarnallite. Collection of Czechoslovak Chemical Communications, 1994, 59, 1620-1623.	1.0	3
53	Isomorphic Co-Crystallization of Ammonium and Rubidium Bromocarnallites. Collection of Czechoslovak Chemical Communications, 1994, 59, 1815-1819.	1.0	5
54	ï»¿Development of accurate chemical thermodynamic database for geochemical storage of nuclear waste. Part II: Models for predicting solution properties and solid-liquid equilibrium in binary nitrate systems. BioRisk, 0, 17, 389-406.	0.2	1

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55	Development of accurate chemical thermodynamic database for geochemical storage of nuclear waste. Part III: Models for predicting solution properties and solid-liquid equilibrium in cesium binary and mixed systems. BioRisk, 0, 17, 407-422.	0.2	1